

To name an ionic compound, state the metal name first, then the nonmetal. Change the ending of the nonmetal to "ide".

Ionic compounds are also called salts. They are always solid, they conduct electricity and they have high melting and boiling points.

Other properties

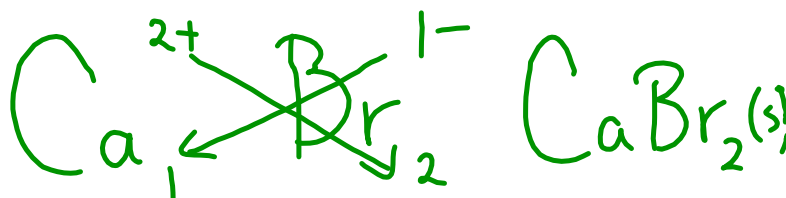
→ brittle

→ have crystal shapes

Writing Formulas for Ionic Compounds

1. Write the symbols, with the metal first.
2. Write the ionic charge above each symbol.
3. Determine how many of each type are needed to balance the charges and make a neutral compound.
4. Write the formula using subscripts to show how many of each ion are needed.

Calcium bromide



Magnesium oxide

copper(II)fluoride

iron(III)oxide

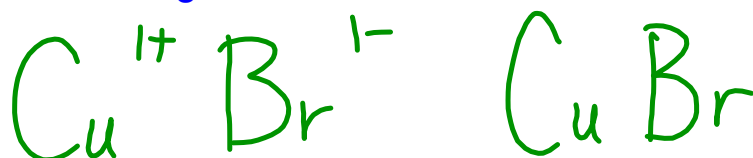
To name ionic compounds with transition metal ions, Roman numerals are placed after the symbol to indicate the ionic charge.

Fe^{+2} is named iron(II)

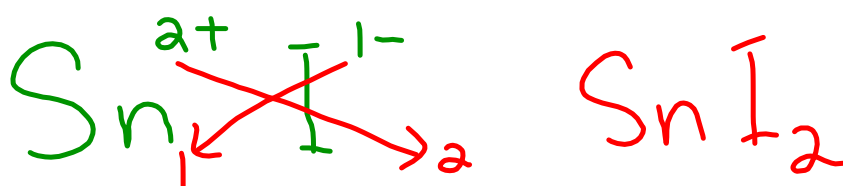
Fe^{+3} is named iron(III)

Writing Formulas

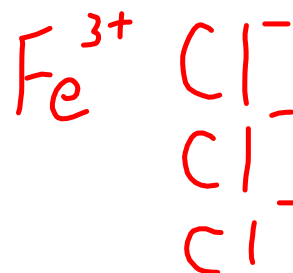
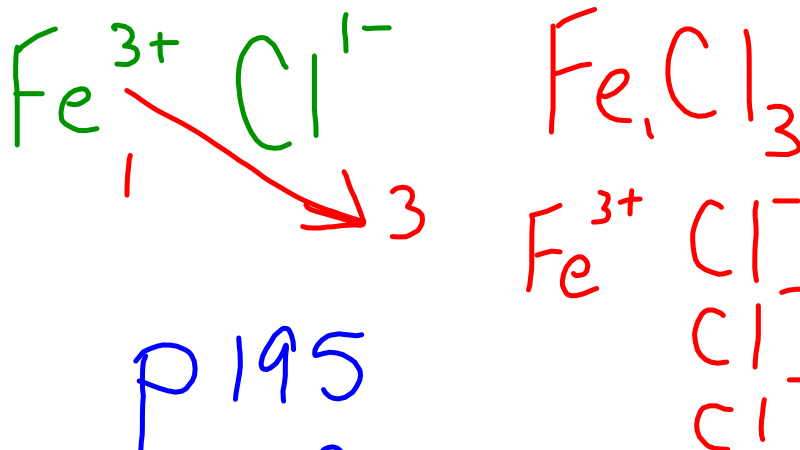
copper(I)bromide



tin(II)iodide



iron(III)chloride



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